

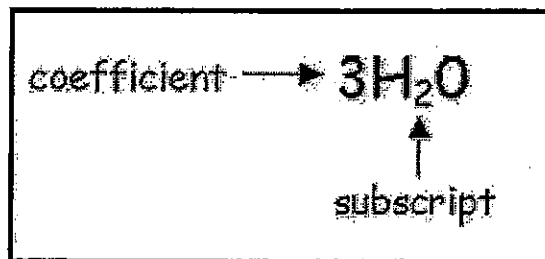
Name \_\_\_\_\_

Date \_\_\_\_\_

## CLASSWORK ~ Practice Counting Atoms Worksheet

Directions for each problem

- 1) write down the different elements in each compound.
- 2) write down how many of that particular atom there are
- 3) how many atoms are there total in the compound.



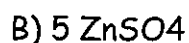
Examples:



Mg - 1

Cl - 2

3 total

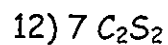
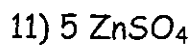
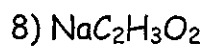
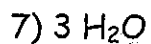
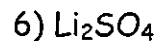


Zn - 5

S - 5

O - 20

30 total



Name \_\_\_\_\_

Date \_\_\_\_\_

Using Science Skills: Making calculations

## Counting Atoms

The formula for a compound indicates the elements that make up the compound and the number of atoms of each element present in the compound. These numbers of atoms are indicated by the use of small numbers called subscripts. Sometimes groups of atoms act as a single atom. Such a group of atoms is called a *polyatomic ion*. If a polyatomic ion is used in a formula more than once, it is put in parentheses and the subscript appears outside the parentheses. When a subscript appears outside the parentheses, it indicates that *all* the elements inside the parentheses should be multiplied by that subscript. For example, the formula  $\text{Fe}(\text{OH})_3$  indicates the combination of one atom of iron, Fe, three atoms of oxygen, O, and three atoms of hydrogen, H.

In the following examples, list each element in the compound and the number of atoms of each element present. The first example has been done for you. You may already be familiar with some of the compounds.

Name	Use	Formula	Atoms in Formula
Calcium carbonate	Limestone	$\text{CaCO}_3$	Ca = calcium 1 C = carbon 1 O = oxygen 3
Aspirin	Pain reliever	$\text{C}_9\text{H}_8\text{O}_4$	
Sucrose	Sugar	$\text{C}_{12}\text{H}_{22}\text{O}_{11}$	
Paradichlorobenzene	Moth crystals	$\text{C}_6\text{H}_4\text{Cl}_2$	
Acetic acid	Found in vinegar	$\text{C}_2\text{H}_4\text{O}_2$	
Silicon dioxide	Sand	$\text{SiO}_2$	
Iron oxide	Rust	$\text{Fe}_2\text{O}_3$	
Pyrite	Fool's gold	$\text{FeS}_2$	